Precipitation Reaction Stoichiometry Practice Problems

1. What mass of sodium chromate is required to precipitate all of the silver ions from 75.0 mL of a 0.100 *M* solution of silver nitrate?
2. What mass of solid silver bromide is produced when 100.0 mL of 0.150 *M* silver nitrate is added to 20.0 mL of 1.00 *M* sodium bromide?
3. A 100.0 mL aliquot of 0.200 *M* of potassium hydroxide is mixed with 100.0 mL of 0.200 *M* magnesium nitrate.
	1. List the species present in each solution
	2. What precipitate forms?
	3. Write a balanced net ionic equation for the reaction.
	4. What mass of precipitate is produced?

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* 1. Calculate the concentration of each ion remaining in solution after precipitation is complete.