Precipitation Reaction Stoichiometry Practice Problems

1. What mass of sodium chromate is required to precipitate all of the silver ions from 75.0 mL of a 0.100 *M* solution of silver nitrate?
2. What mass of solid silver bromide is produced when 100.0 mL of 0.150 *M* silver nitrate is added to 20.0 mL of 1.00 *M* sodium bromide?
3. A 100.0 mL aliquot of 0.200 *M* of potassium hydroxide is mixed with 100.0 mL of 0.200 *M* magnesium nitrate.
   1. List the species present in each solution
   2. What precipitate forms?
   3. Write a balanced net ionic equation for the reaction.
   4. What mass of precipitate is produced?

|  |  |  |  |
| --- | --- | --- | --- |
|  |  |  |  |
| **I** |  |  |  |
| **C** |  |  |  |
| **E** |  |  |  |

* 1. Calculate the concentration of each ion remaining in solution after precipitation is complete.